

0,6000 mol



5 mol C

10 mol H

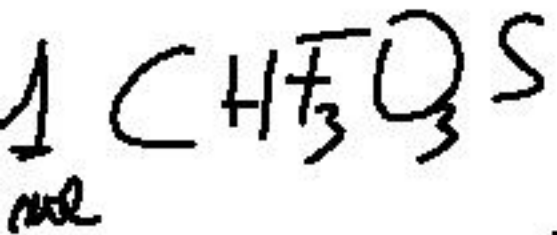
5 mol O

$$1 \text{ mol}_{C_5 H_{10} O_3} : 5 \text{ mol}_C = 0,6000 \text{ mol} \cdot X$$

$$X = 3 \text{ mol}$$

$$N_{\text{atom C}} = 3 \text{ mol} \cdot 6,022 \cdot 10^{23} \frac{\text{atom}}{\text{mol}} = 18,066 \cdot 10^{23} \text{ atom C}$$

$$4,62 \cdot 10^{23} \text{ atomi O}$$



$$6,022 \cdot 10^{23} \text{ atomi : 1} = 4,62 \cdot 10^{23} : X$$

$$X = \frac{4,62 \cdot 10^{23} \text{ atomi O}}{6,022 \cdot 10^{23} \text{ atomi} \cdot \text{mol}^{-1}} = 0,76 \text{ mol O}$$

Rapporto molare

Composto : O

$$1 : 3 = X : 0,76 \text{ molo}$$

$$X = 0,76 \text{ mol} / 3 = 0,25 \text{ mol}$$

