

$Ca_3(PO_4)_2$ da 22g di K_3PO_4

12g di $CaCl_2$

$$MM = (39,10 \cdot 3) + (30,97) + (16 \cdot 4) \text{ g} \cdot \text{mol}^{-1}$$
$$117,3 + 30,97 + 64 = 212,27 \text{ g} \cdot \text{mol}^{-1}$$

$$n_{K_3PO_4} = \frac{22g}{212,27 \text{ g} \cdot \text{mol}^{-1}} = 0,10 \text{ mol}$$

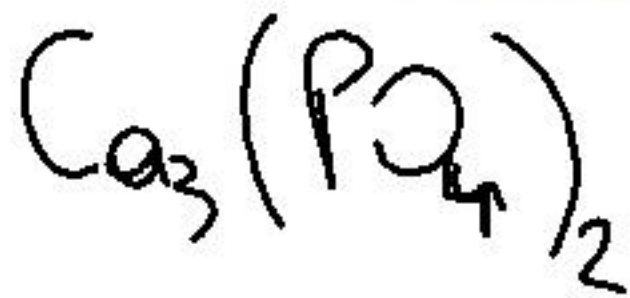
$$2:3 = 0,10 : x$$

$x = 0,15 \text{ mol}$ di $CaCl_2$

$$MM_{CaCl_2} = 40,08 + (35,45 \cdot 2) \text{ g} \cdot \text{mol}^{-1}$$
$$= 110,98 \text{ g} \cdot \text{mol}^{-1}$$

$$n_{\text{CaCl}_2} = \frac{12 \text{ g}}{110,98 \text{ g mol}^{-1}} = 0,10 \text{ mol}$$

CaCl_2 limitante



$$x = 0,10 : 3 = 0,0\bar{3} \text{ mol di } \text{Ca}_3(\text{PO}_4)_2$$

$$\text{Massa} = 0,03 \text{ mol} \cdot [120 + (31 + 64) \cdot 2] \text{ g mol}^{-1} = 12,9 \text{ g}$$

